

BALANCING CHEMICAL EQUATIONS PRACTICE

1. $1 \text{ H}_2 + 1 \text{ O} \rightarrow 1 \text{ H}_2\text{O}$	15. $2 \text{ Al} + 3 \text{ Br}_2 \rightarrow 2 \text{ AlBr}_3$
2. $3 \text{ H}_2 + 1 \text{ N}_2 \rightarrow 2 \text{ NH}_3$	16. $4 \text{ Cr} + 3 \text{ O}_2 \rightarrow 2 \text{ Cr}_2\text{O}_3$
3. $2 \text{ Fe} + 3 \text{ Cl}_2 \rightarrow 2 \text{ Fe}_2\text{Cl}_3$	17. $2 \text{ NaClO}_3 \rightarrow 2 \text{ NaCl} + 3 \text{ O}_2$
4. $4 \text{ Fe} + 3 \text{ O}_2 \rightarrow 2 \text{ Fe}_2\text{O}_3$	18. $2 \text{ AlBr}_3 + 3 \text{ Cl}_2 \rightarrow 2 \text{ AlCl}_3 + 3 \text{ Br}_2$
5. $2 \text{ Al}_2\text{O}_3 \rightarrow 4 \text{ Al} + 3 \text{ O}_2$	19. $1 \text{ C}_4\text{H}_{10}\text{O} + 6 \text{ O}_2 \rightarrow 4 \text{ CO}_2 + 5 \text{ H}_2\text{O}$
6. $2 \text{ KClO}_3 \rightarrow 2 \text{ KCl} + 3 \text{ O}_2$	20. $1 \text{ C}_7\text{H}_{16} + 11 \text{ O}_2 \rightarrow 7 \text{ CO}_2 + 8 \text{ H}_2\text{O}$
7. $1 \text{ S}_8 + 8 \text{ O}_2 \rightarrow 8 \text{ SO}_2$	21. $2 \text{ Na} + 2 \text{ H}_2\text{O} \rightarrow 2 \text{ NaOH} + 1 \text{ H}_2$
8. $2 \text{ FeBr}_3 + 3 \text{ H}_2\text{SO}_4 \rightarrow 1 \text{ Fe}_2(\text{SO}_4)_3 + 6 \text{ HBr}$	22. $2 \text{ AlI}_3 + 3 \text{ HgCl}_2 \rightarrow 2 \text{ AlCl}_3 + 3 \text{ HgI}_2$
9. $2 \text{ C}_2\text{H}_6 + 7 \text{ O}_2 \rightarrow 4 \text{ CO}_2 + 6 \text{ H}_2\text{O}$	23. $3 \text{ Ca}(\text{OH})_2 + 2 \text{ H}_3\text{PO}_4 \rightarrow 1 \text{ Ca}_3(\text{PO}_4)_2 + 6 \text{ H}_2\text{O}$
10. $1 \text{ C}_4\text{H}_6\text{O}_3 + 1 \text{ H}_2\text{O} \rightarrow 2 \text{ C}_2\text{H}_4\text{O}_2$	24. $4 \text{ H}_2\text{SiCl}_2 + 4 \text{ H}_2\text{O} \rightarrow 1 \text{ H}_8\text{Si}_4\text{O}_4 + 8 \text{ HCl}$
11. $1 \text{ C}_2\text{H}_4 + 3 \text{ O}_2 \rightarrow 2 \text{ CO}_2 + 2 \text{ H}_2\text{O}$	25. $10 \text{ HSiCl}_3 + 15 \text{ H}_2\text{O} \rightarrow 1 \text{ H}_{10}\text{Si}_{10}\text{O}_{15} + 30 \text{ HCl}$
12. $1 \text{ Al}_2(\text{SO}_4)_3 + 3 \text{ Ca}(\text{OH})_2 \rightarrow 2 \text{ Al}(\text{OH})_3 + 3 \text{ CaSO}_4$	26. $3 \text{ AgNO}_3 + 1 \text{ K}_3\text{PO}_4 \rightarrow 1 \text{ Ag}_3\text{PO}_4 + 3 \text{ KNO}_3$
13. $1 \text{ P}_4 + 5 \text{ O}_2 \rightarrow 2 \text{ P}_2\text{O}_5$	27. $1 \text{ C}_7\text{H}_9 + 3 \text{ HNO}_3 \rightarrow 1 \text{ C}_7\text{H}_6(\text{NO}_2)_3 + 3 \text{ H}_2\text{O}$
14. $16 \text{ Ag} + 1 \text{ S}_8 \rightarrow 8 \text{ Ag}_2\text{S}$	28. $1 \text{ C}_5\text{H}_8\text{O}_2 + 2 \text{ NaH} + 2 \text{ HCl} \rightarrow 1 \text{ C}_5\text{H}_{12}\text{O}_2 + 2 \text{ NaCl}$